



## Cambridge IGCSE™

## CO-ORDINATED SCIENCES

0654/13

Paper 1 Multiple Choice (Core)

October/November 2023

45 minutes

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet  
Soft clean eraser  
Soft pencil (type B or HB is recommended)

## INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A**, **B**, **C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

## INFORMATION

- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.

This document has **16** pages.



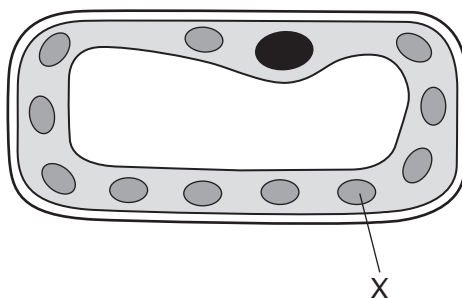
## 2

- 1 A person throws a ball which their dog runs after and brings back to them.

Which characteristics of living things is the dog showing by this action?

- A growth and nutrition
- B movement and nutrition
- C movement and sensitivity
- D sensitivity and growth

- 2 The diagram shows a palisade cell from a leaf magnified  $\times 250$ .



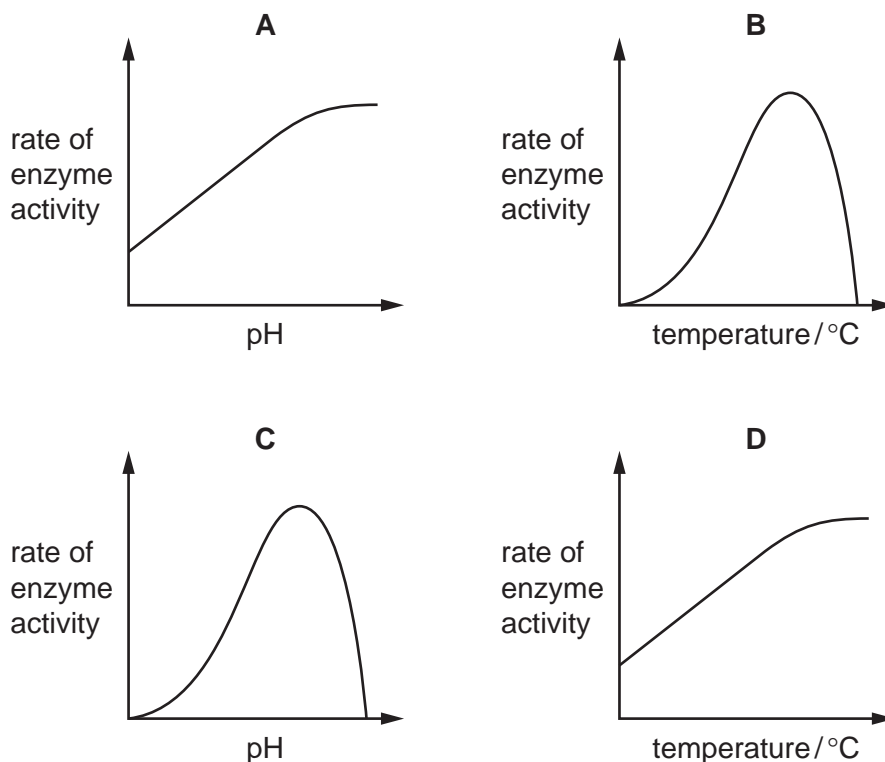
The chloroplast labelled X measures 5 mm on the diagram.

What is its actual length?

- A 0.25 mm
  - B 0.05 mm
  - C 0.02 mm
  - D 0.01 mm
- 3 Which food test requires heating?
- A fat
  - B protein
  - C reducing sugar
  - D starch

3

4 Which graph is correct?



5 Plants require various ions to manufacture other molecules.

Which row shows the molecules that plants make from magnesium ions and nitrate ions?

	magnesium ions	nitrate ions
<b>A</b>	chlorophyll	chlorophyll
<b>B</b>	chlorophyll	protein
<b>C</b>	protein	chlorophyll
<b>D</b>	protein	protein

6 What is a good source of vitamin C?

- A** citrus fruits
- B** liver
- C** meat
- D** milk

## 4

7 Which row names the substances carried by xylem vessels and the direction of travel?

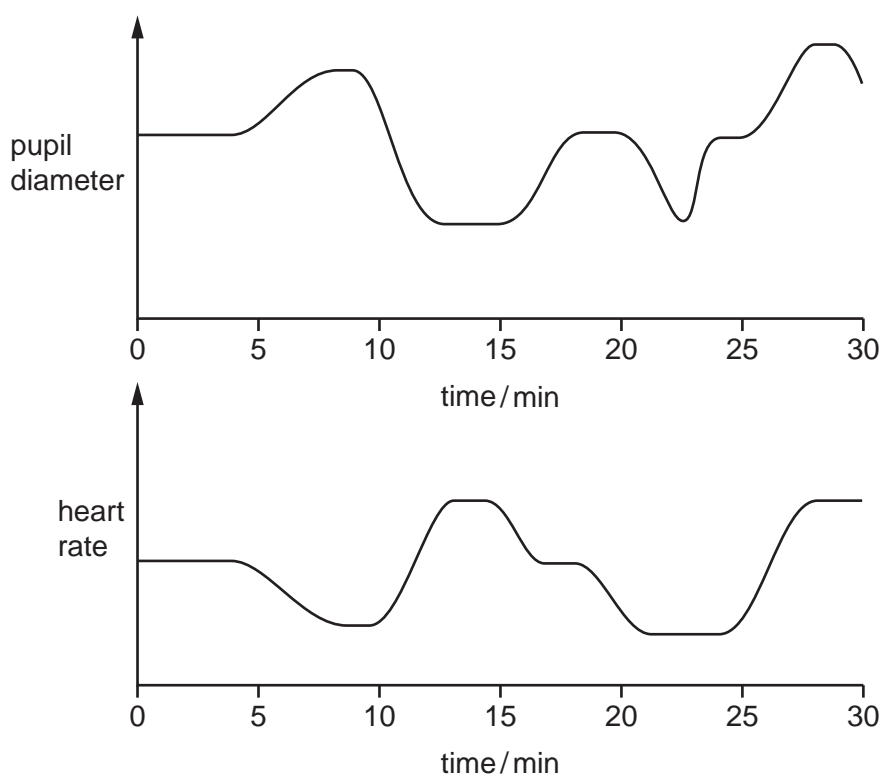
	substances	direction of travel
<b>A</b>	sucrose only	leaves to roots
<b>B</b>	sucrose only	roots to leaves
<b>C</b>	water and dissolved minerals	leaves to roots
<b>D</b>	water and dissolved minerals	roots to leaves

8 What is the expected concentration of oxygen and the water vapour content in expired air?

	oxygen / %	water vapour
<b>A</b>	16	saturated
<b>B</b>	16	variable
<b>C</b>	21	saturated
<b>D</b>	21	variable

5

- 9 The graphs show the pupil diameter and heart rate of a person over a period of 30 minutes. The person was given an unexpected shock at one point during the 30 minutes.



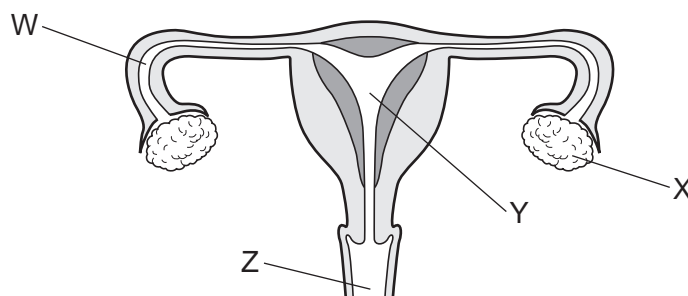
At which time did the person experience a shock causing them to release adrenaline?

- A 5 minutes
- B 10 minutes
- C 15 minutes
- D 25 minutes

**10** There are different conditions which may prevent a successful pregnancy.

- 1 failure of ovulation
- 2 failure of embryo to implant into uterus lining
- 3 blockage of oviducts preventing fertilisation

The diagram shows the human female reproductive system.



Which locations are involved in each condition?

	1	2	3
<b>A</b>	W	X	Z
<b>B</b>	W	Y	X
<b>C</b>	X	Y	W
<b>D</b>	X	Z	W

**11** Which combination of sex chromosomes are present in the body cells of a human male?

- A** XX                      **B** XY                      **C** Y only                      **D** YY

**12** The diagram shows a food chain.

mahogany tree → caterpillar → songbird → hawk

What is the mahogany tree in this food chain?

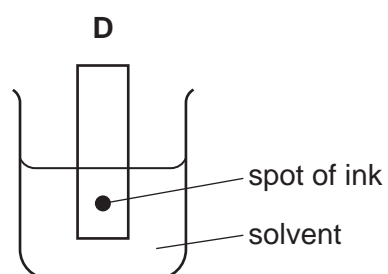
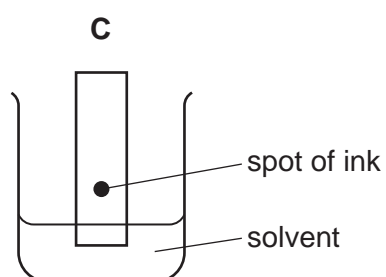
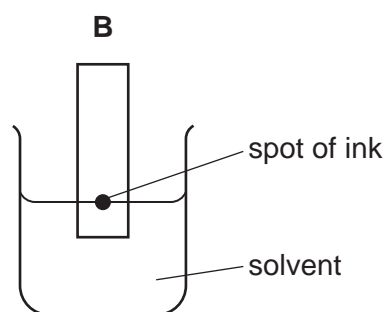
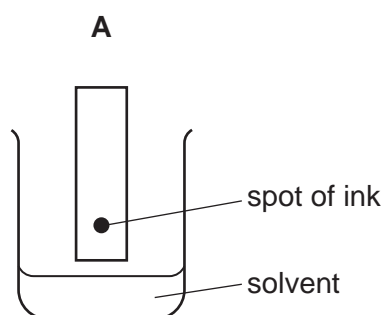
- A** carnivore  
**B** consumer  
**C** herbivore  
**D** producer

13 Which row describes the effects of deforestation?

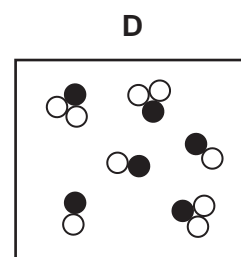
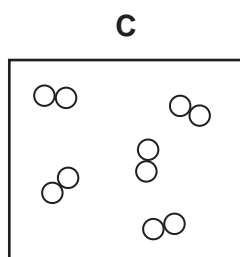
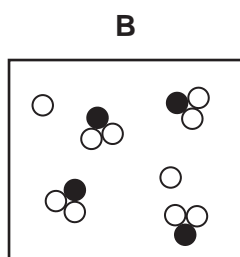
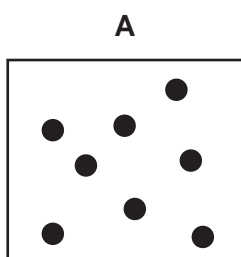
	level of carbon dioxide in the air	risk of flooding
<b>A</b>	decreases	decreases
<b>B</b>	decreases	increases
<b>C</b>	increases	decreases
<b>D</b>	increases	increases

14 The colours in an ink are separated by chromatography.

Which diagram shows the assembled apparatus?



15 Which diagram represents a mixture of an element and a compound?



- 16 X and Y are isotopes of the same element.

Which statement about X and Y is correct?

- A** They have the same nucleon number but different numbers of protons.
- B** They have the same number of neutrons but different numbers of electrons.
- C** They have the same atomic number but different numbers of electrons.
- D** They have the same number of protons but different numbers of neutrons.
- 17 Dilute hydrochloric acid reacts with sodium carbonate to form sodium chloride, carbon dioxide and water.

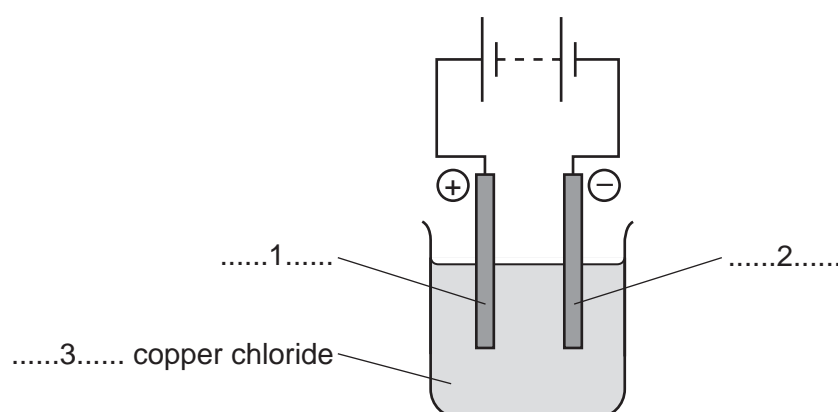
The equation for the reaction is shown.



What are the values of  $w$ ,  $x$ ,  $y$  and  $z$ ?

	$w$	$x$	$y$	$z$
<b>A</b>	1	1	3	1
<b>B</b>	2	1	1	1
<b>C</b>	2	2	3	2
<b>D</b>	2	2	1	1

- 18 The apparatus used for the electrolysis of copper chloride is shown.

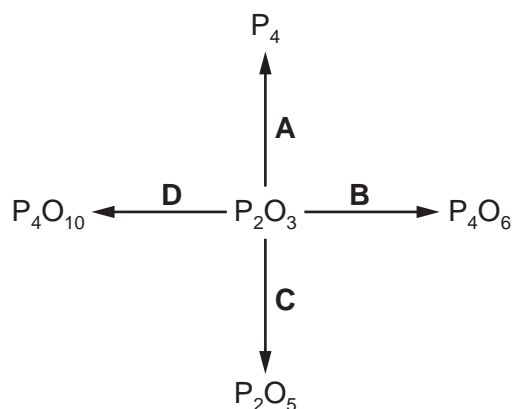


Which words complete labels 1, 2 and 3?

	1	2	3
<b>A</b>	anode	cathode	aqueous
<b>B</b>	anode	cathode	solid
<b>C</b>	cathode	anode	aqueous
<b>D</b>	cathode	anode	solid



- 19 In which change is the oxide of phosphorus,  $P_2O_3$ , reduced?



- 20 A piece of damp blue litmus paper is put in a test-tube of a gas. The litmus paper turns red and then changes to white.

What is the gas?

- A** ammonia
  - B** carbon dioxide
  - C** chlorine
  - D** oxygen
- 21 Element X burns in air to produce an oxide.

A solution of the oxide has a pH value of 10.

What is X?

- A** argon
  - B** carbon
  - C** magnesium
  - D** phosphorus
- 22 The elements in Group I of the Periodic Table are metals.

What are the trends as the group is descended?

- A** decrease in melting point and less reactive with water
- B** decrease in melting point and more reactive with water
- C** increase in melting point and less reactive with water
- D** increase in melting point and more reactive with water

**23** Why is argon used in lamps?

- A** It is heavier than air.
- B** It is lighter than air.
- C** It is reactive.
- D** It is unreactive.

**24** Which statement explains why carbon is used in the extraction of copper from its ore?

- A** It is in Group IV of the Periodic Table.
- B** It is more reactive than copper.
- C** It is a non-metal.
- D** It forms a giant covalent molecule.

**25** Which substances are needed for iron to rust?

- A** oxygen and water
- B** nitrogen and water
- C** oxygen only
- D** water only

**26** Which energy source is an example of a fossil fuel?

- A** hydrogen
- B** methane
- C** the Sun
- D** wood

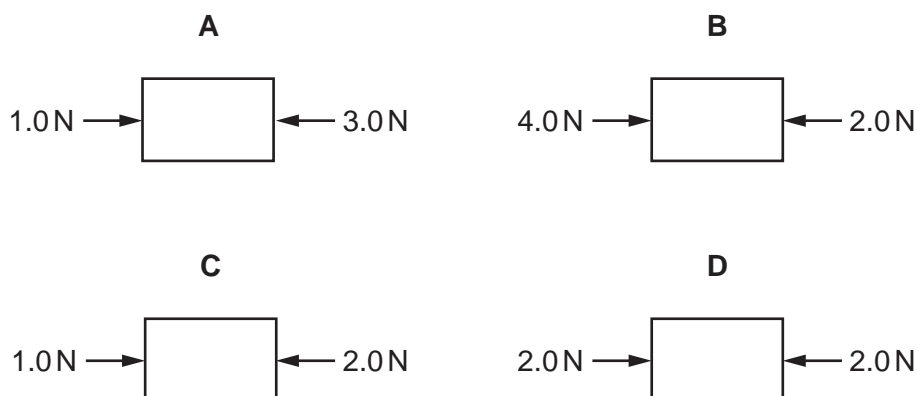
**27** Molecules of ethene react together to form long chain molecules.

What is the name of this type of reaction?

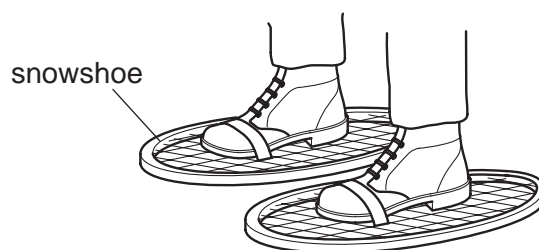
- A** addition polymerisation
- B** cracking
- C** fermentation
- D** thermal decomposition

28 The diagrams show four objects, each acted on by only two forces.

Which object is in equilibrium?



29 A person wearing snowshoes is walking on snow.



How do snowshoes decrease the pressure on the snow?

- A** by applying the person's weight over a larger area
- B** by applying the person's weight over a smaller area
- C** by decreasing the person's weight
- D** by increasing the person's weight

30 A machine does a certain quantity of work in a certain time.

Which changes to the quantity of work done and to the time taken **both** decrease the power produced by the machine?

	work done	time taken
<b>A</b>	decrease	decrease
<b>B</b>	decrease	increase
<b>C</b>	increase	decrease
<b>D</b>	increase	increase

- 31** A sample of liquid cools for 20 minutes. Its temperature is recorded every 2 minutes.

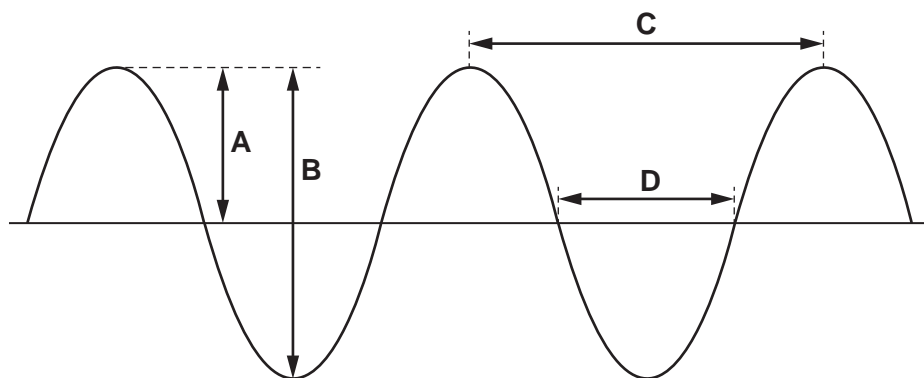
The results are shown.

time / minutes	0	2	4	6	8	10	12	14	16	18	20
temperature / °C	90.8	80.9	74.1	67.4	61.9	57.0	53.0	50.2	48.5	47.3	46.1

How should the sample be described at the end of 18 minutes?

- A** all liquid
- B** all solid
- C** in the process of boiling
- D** in the process of solidifying
- 32** In which of the three states of matter is convection the main method of thermal energy transfer?
- A** liquids and gases
- B** liquids only
- C** solids and liquids
- D** solids only
- 33** The diagram represents a transverse wave.

Which labelled arrow represents the wavelength of the wave?

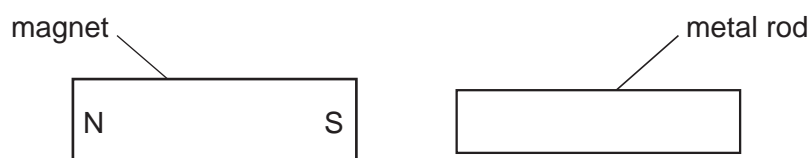


- 34** The sound heard from the siren of a police car becomes quieter and lower pitched as the car moves away from an observer.

Which row describes what happens to the amplitude and frequency of the sound wave heard by the observer?

	amplitude	frequency
<b>A</b>	decreases	decreases
<b>B</b>	decreases	increases
<b>C</b>	increases	decreases
<b>D</b>	increases	increases

- 35** A bar magnet is brought near to a metal rod. The metal rod is attracted to the magnet.



The magnet is then turned around so that the N-pole is on the right.

The magnet is again brought near to the metal rod and is again attracted to the magnet.

What could the metal rod be?

- A** another bar magnet
  - B** a piece of aluminium
  - C** a piece of copper
  - D** a piece of iron
- 36** Four resistors are connected into circuits. The current in each resistor and the potential difference (p.d.) across each resistor are shown.

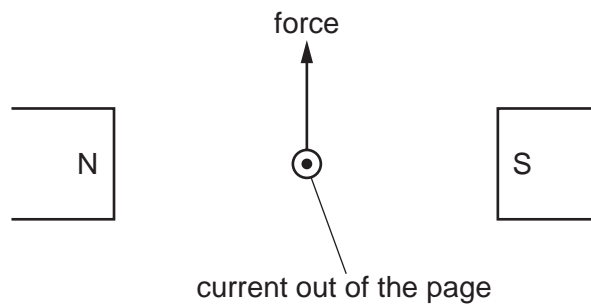
Which resistor has a resistance of  $2.0\ \Omega$ ?

	current / A	p.d. / V
<b>A</b>	2.0	1.0
<b>B</b>	4.0	2.0
<b>C</b>	12	6.0
<b>D</b>	4.0	8.0

- 37 Two lamps can be connected to a battery either in series or in parallel.

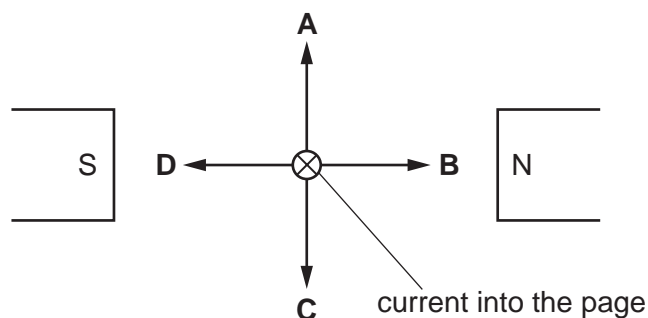
Which statement is **not** a benefit of connecting two lamps in parallel rather than in series?

- A If one lamp breaks, the other lamp stays lit.
  - B The lamps are brighter.
  - C The lamps can be controlled individually using switches.
  - D There is a smaller current in the battery.
- 38 What is the purpose of a fuse in an electric circuit?
- A to make the circuit more efficient
  - B to protect the circuit from damage by a large current
  - C to provide a constant current in the circuit
  - D to provide a constant potential difference (p.d.) across the circuit
- 39 The diagram shows a current-carrying wire placed between the poles of a magnet. The direction of the current is out of the page. The direction of the force on the wire is shown by the arrow.



Both the direction of the current and the poles of the magnet are now reversed.

Which arrow shows the direction of the force on the wire after these changes?



40 The nucleus of an atom emits an  $\alpha$ -particle.

How do the number of protons and the number of neutrons in the nucleus change?

	number of protons	number of neutrons
<b>A</b>	decreases by 2	decreases by 2
<b>B</b>	decreases by 2	decreases by 4
<b>C</b>	increases by 2	increases by 2
<b>D</b>	increases by 2	increases by 4

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The Periodic Table of Elements

Group																	
I	II											III	IV	V	VI	VII	VIII
<div>1Hhydrogen1</div> <div>Key<div>atomic number atomic symbol name relative atomic mass</div></div>																	
3Li lithium 7	4Be beryllium 9											5B boron 11	6C carbon 12	7N nitrogen 14	8O oxygen 16	9F fluorine 19	10Ne neon 20
11Na sodium 23	12Mg magnesium 24											13Al aluminium 27	14Si silicon 28	15P phosphorus 31	16S sulfur 32	17Cl chlorine 35.5	18Ar argon 40
19K potassium 39	20Ca calcium 40	21Sc scandium 45	22Ti titanium 48	23V vanadium 51	24Cr chromium 52	25Mn manganese 55	26Fe iron 56	27Co cobalt 59	28Ni nickel 59	29Cu copper 64	30Zn zinc 65	31Ga gallium 70	32Ge germanium 73	33As arsenic 75	34Se selenium 79	35Br bromine 80	36Kr krypton 84
37Rb rubidium 85	38Sr strontium 88	39Y yttrium 89	40Zr zirconium 91	41Nb niobium 93	42Mo molybdenum 96	43Tc technetium —	44Ru ruthenium 101	45Rh rhodium 103	46Pd palladium 106	47Ag silver 108	48Cd cadmium 112	49In indium 115	50Sn tin 119	51Sb antimony 122	52Te tellurium 128	53I iodine 127	54Xe xenon 131
55Cs caesium 133	56Ba barium 137	57–71 lanthanoids	72Hf hafnium 178	73Ta tantalum 181	74W tungsten 184	75Re rhenium 186	76Os osmium 190	77Ir iridium 192	78Pt platinum 195	79Au gold 197	80Hg mercury 201	81Tl thallium 204	82Pb lead 207	83Bi bismuth 209	84Po polonium —	85At astatine —	86Rn radon —
87Fr francium —	88Ra radium —	89–103 actinoids	104Rf rutherfordium —	105Db dubnium —	106Sg seaborgium —	107Bh bohrium —	108Hs hassium —	109Mt meitnerium —	110Ds darmstadtium —	111Rg roentgenium —	112Cn copernicium —	113Nh nihonium —	114Fl flerovium —	115Mc moscovium —	116Lv livermorium —	117Ts tennessine —	118Og oganeson —

lanthanoids		57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
actinoids		89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

The volume of one mole of any gas is 24dm<sup>3</sup> at room temperature and pressure (r.t.p.).